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Divisions: AMERICAN CHEMICAL SOCIETY; American
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Engineering (See All Technical Divisions) ISBN13:
9780841202009. eISBN: 9780841223141. DOI:
10.1021/ba-1974-0133.

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Solve problem 1-15 from Elements of Chemical Reaction Engineering.

Book Problem 1-15 (Elements of Chemical Reaction Engineering)

This page contains lecture notes from a typical Chemical Reaction Engineering class. Two different sources of lecture notes are provided from the respective professors and their institutions. University of Michigan: University of Illinois: Che344- Chemical Reactions Engineering:

Elements of Chemical Reaction Engineering

Chemical Reaction Engineering II. L1-Introduction to catalysts and catalysis; L2-Steps in catalytic reaction: adsorption, desorption and reaction

NPTEL :: Chemical Engineering - Chemical Reaction ...

For 30 years, H. Scott Fogler's Elements of Chemical Reaction Engineering has been the #1 selling text for courses in chemical reaction engineering worldwide. Now, in Essentials of Chemical Reaction Engineering, Second Edition, Fogler has distilled this classic into a modern, introductory-level guide specifically for undergraduates.

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Exchange Equipments - Evaporation - Diffusion - Distillation - Gas Absorption - Liquid Liquid Extraction - Crystallisation - Drying - Appendix I Try yourself - Appendix II Thermal conductivity data - Appendix III Steam tables

Chemical reaction engineering is at the core of chemical engineering education. Unfortunately, the subject can be intimidating to students, because it requires a heavy dose of mathematics. These mathematics, unless suitably explained in the context of the physical phenomenon, can confuse rather than enlighten students. Bearing this in mind, Reaction Engineering Principles is written primarily from a student's perspective. It is the culmination of the author's more than twenty years of experience teaching chemical reaction engineering. The textbook begins by covering the basic building blocks of the subject—stoichiometry, kinetics, and thermodynamics—ensuring students gain a good grasp of the essential concepts before venturing into the world of reactors. The design and performance evaluation of reactors are conveniently grouped into chapters based on an increasing degree of difficulty. Accordingly, isothermal reactors—batch and ideal flow types—are addressed first, followed by non-isothermal reactor operation, non-ideal flow in reactors, and some special reactor types. For better comprehension, detailed derivations are provided for all important mathematical equations. Narrative of the physical context in which the formulae work adds to the clarity of thought. The use of mathematical formulae is elaborated upon in the form of problem solving steps followed by worked examples. Effects of parameters, changing trends, and comparisons between different situations are presented graphically. Self-practice exercises are included at the end of each chapter.

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Fuels and combustion. Gas producers. Sulfur compounds. Metallurgy. Crystallization.

The role of the chemical reactor is crucial for the industrial conversion of raw materials into products and numerous factors must be considered when selecting an appropriate and efficient chemical reactor. Chemical Reaction Engineering and Reactor Technology defines the qualitative aspects that affect the selection of an industrial chemical reactor and couples various reactor models to case-specific kinetic expressions for chemical processes. Thoroughly revised and updated, this much-anticipated Second Edition addresses the rapid academic and industrial development of chemical reaction engineering. Offering a systematic development of the chemical reaction engineering concept, this volume explores: essential stoichiometric, kinetic, and thermodynamic terms needed in the analysis of chemical reactors homogeneous and heterogeneous reactors reactor optimization aspects residence time distributions and non-ideal flow conditions in industrial reactors solutions of algebraic and ordinary differential equation systems gas- and liquid-phase diffusion coefficients and gas-film coefficients correlations for gas-liquid systems solubilities of gases in liquids guidelines for laboratory reactors and the estimation of kinetic parameters The authors pay special attention to the exact formulations and derivations of mass energy balances and their numerical solutions. Richly illustrated and containing exercises and solutions covering a number of processes, from oil refining to the development of specialty and fine chemicals, the text provides a clear understanding of chemical reactor analysis and design.

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Designed as an undergraduate-level textbook in Chemical Engineering, this student-friendly, thoroughly class-room tested book, now in its second edition, continues to provide an in-depth analysis of chemical engineering thermodynamics. The book has been so organized that it gives comprehensive coverage of basic concepts and applications of the laws of thermodynamics in the initial chapters, while the later chapters focus at length on important areas of study falling under the realm of chemical thermodynamics. The reader is thus introduced to a thorough analysis of the fundamental laws of thermodynamics as well as their applications to practical situations. This is followed by a detailed discussion on relationships among thermodynamic properties and an exhaustive treatment on the thermodynamic properties of solutions. The role of phase equilibrium thermodynamics in design, analysis, and operation of chemical separation methods is also deftly dealt with. Finally, the chemical reaction equilibria are skillfully explained. Besides numerous illustrations, the book contains over 200 worked examples, over 400 exercise problems (all with answers) and several objective-type questions, which enable students to gain an in-depth understanding of the concepts and theory discussed. The book will also be a useful text for students pursuing courses in chemical engineering-related branches such as polymer engineering, petroleum engineering, and safety and environmental engineering. New to This Edition

- More Example Problems and Exercise Questions in each chapter
- Updated section on Vapour-Liquid Equilibrium in Chapter 8 to highlight the significance of equations of state approach
- GATE Questions up to 2012 with answers

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